5		0/05/2016		
ſime	bject Name: Pharmaceutical Analysis - I me: 2:30 PM to 5:30 PM Total M		arks: 80	
1. 2.	Make s	ot any five questions. uitable assumptions wherever necessary. s to the right indicate full marks.		
Q.1	(a)	What is hydrolysis? Derive equation for finding pH of aqueous	06	
	(b)	solution of salt of strong acid and weak base. Write a note on pharmaceutical buffers.	05	
	(b) (c)	At what molar concentration Acetic acid is 2% ionized?	05	
	(0)	(Ka=8 X 10^{-5})	05	
Q.2	(a)	Explain the terms:	06	
		(i) Co-precipitation (ii) Post precipitation (iii) Solubility product		
	(b)	What is buffer? Derive Henderson-Hasselbach equation.	05	
	(c)	Write a note on acid base indicator theory.	05	
Q.3	(a)	What is Argentometric titration? Write a note on Mohrs method.	06	
	(b)	Ksp of PbI ₂ is 7.2 X 10^{-9} . Calculate molar solubility and solubility in g/ml. Molecular weight of PbI ₂ is 461 g/mol.	05	
	(c)	Write a note on Oxygen combustion flask method.	05	
Q.4	(a)	Write theory of complexometric titration? Masking and damasking of complexometric titration.	06	
	(b)	Explain indicators in complexometric titration?	05	
	(c)	Explain Kjeldahl method for determination of nitrogen.	05	
Q.5	(a)	What is non-aqueous titration? Write Differentiating and leveling effect of solvent.	06	
	(b)	Discuss Merits and Demerits of non aqueous titration over aqueous titration.	05	
	(c)	Write a note on Karl-Fischer titration.	05	
Q. 6	(a)	Write a note on Diazotization nitrite titration.	06	
	(b)	Differentiate Iodometric and Iodimetric titrations.	05	
	(c)	Give preparation and standardization of KMnO ₄	05	
Q.7	(a)	What is analytical method validation? Enlist validation parameter. Discuss any two parameters in detail.	06	
	(b)	What is error? Write types of errors and how they minimized.	05	
	(c)	Differentiate between QA and QC.	05	